Human Gas Demo HW

Read and outline (Cornell style) **Section 14.1** in your chemistry textbooks. ****Your section outline must be at least ½ page with 3-4 topic questions.**** Then answer the following assessment questions:

1. State Boyle's, Charles', and Gay-Lussac's laws using sentences, then equations.
2. Convert the following pressures:
    a. 1.75 atm = _____ torr
    b. 2.04 atm = _____ kPa
    c. 0.21 atm = _____ psi
    d. 973 torr = _____ mm Hg
    e. 187 torr = _____ atm
3. The air pressure for a certain tire is 109 kPa. What is this pressure in atmospheres?
4. The air pressure inside a submarine is 0.62 atm. What would be the height of a column of mercury balanced by this pressure?
5. The weather news gives the atmospheric pressure as 1.07 atm. What is this atmospheric pressure in torr?